  of and of solution are mixed. Molarity of  ions in the resulting solution is: (At. wt. of )

A mixture of  and gas “” (molecular weight ) in the mole ratio  has a mean molecular weight . What would be mean molecular weight, if the gases are mixed in the ratio  under identical conditions? (gases are non-reacting):

A mixture of  and contain  mass per cent of nitrogen. What is the mass ratio of the two components in the mixture? (Atomic mass of )

Suppose two elements  and  combine to form two compounds  and . When  of  weighs   while  molecules of  weighs  . The atomic masses of  and  are respectively:

  of a sample on complete combustion gives   and of . The molecular formula of the compound may be:

Which statement is false for the balanced equation given below?

A air sample contains: Oxygen  Carbon  Hydrogen , and Nitrogen  The weight which the air sample would gain if all  atoms are replaced by  atoms is:

What is the empirical formula of vanadium oxide, if   of the metal oxide contains   of metal? (Atomic mass of )

The ratio of masses of oxygen and nitrogen in a particular gaseous mixture is . The ratio of number of their molecules is:

   and    are mixed in volume ratio of . What is the final molarity of  solution?

Molar mass of an unknown compound is found to be $46 \; &nbsp;g/mol$. What will be its molecular mass?

â???~If two elements can combine to form more than one compound, the masses of one element that combine with a fixed mass of the other element are in the ratio of small whole numbers.â???T This statement states:

Barium chloride and sodium phosphate react according to the equation

If $0.5 &nbsp;mole$ of is mixed with $0.2\; &nbsp;mole$ of , how many  moles of are formed?

A mixture of and is caused to react in a closed container to form . The reaction ends before either reactant has been totally consumed. At this stage, each of and are present. The moles of and present originally were respectively

of menthol () burns in air to give of and $1.35\; &nbsp;g$ of . What is the empirical formula of menthol, if the combustion reaction can be represented as:

A gaseous mixture contains and ) in a ratio by mass. The ratio of the number of molecules of and is

On heating pure limestone dissociates as . If $200\; &nbsp;g$ of limestone is heated in an open container, compute the reduction in its weight. ( Atomic mass of $Ca = 40, C = 12, O = &nbsp;16$)

of a substance of molecular weight when dissolved in  water yields a solution of density $1.1\; &nbsp;g / mL$. The molarity of the solution is

Consider  the reaction . In which of the following reaction mixtures is the limiting reagent?

Which of the following statements is incorrect?